

Chemistry B Moles Packet Answer Key

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CHEMISTRY WORKSHEET # 6 MIXED MOLE PROBLEMS (GRAMS, MOLECULES, AND LITERS) You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4

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liters at STP.

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CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS

Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.

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While a dozen is only 12 particles a mole is a much larger number— 6.02×10^{23} particles. Elements generally exist as the particles we call atoms. A mole of carbon contains 6.02×10^{23} atoms of carbon. A mole of helium contains 6.02×10^{23} atoms of helium. A mole of sodium contains 6.02×10^{23} atoms of sodium.

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Chemistry Stoichiometry Packet Answers

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INTRODUCTION TO MOLES. Molar Mass (# grams) 6.02×10^{23}
ATOMS (for elements) or $23 \times 6.02 \times 10^{23}$ MOLECULES (for
compounds) Atomic Weight (for elements) or Gram
Formula/Molecular Weight (for compounds) MOLES 22.4 LITERS
OF A GAS AT STP (standard temperature and pressure) FORMULA
CALCULATION MASS CHEMISTRY WORKSHEET # 2: THE MOLE AS
A UNIT OF ...

chemistry worksheet # 2: the mole as a unit of mass

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4CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF

MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique. 1 mol = a molar mass of an atom/molecule

(g/mol) EXAMPLE: What is the mass of 5.00 moles of water H ...

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A) 8.5 moles of sodium acetate dissolved in 1 L of water B) 5.5 moles of sodium acetate dissolved in 500 mL of water C) 1.8 moles of sodium acetate dissolved in 300 mL of water D) 1.2 moles of sodium acetate dissolved in 200 mL of water E) none of the above

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